	Chemical Reactions Practice	Quiz #2	Name	D	
\ <u>!</u>	Write the balanced equations for each Write the net ionic equation for all dou	of the following	1 " 1		n with the
t	<i>Write the net ionic equation for all dou</i> he space provided. Be sure to include	<i>ble displacemen</i> states of matte	t reactions that occur	. Write the types of	reaction in
	olutions of potassium carbonate and ni		•		
2) So	odium metal is placed in water				
3) Co	pper is placed in a solution of nickel (II)	nitrate. (The co	pper (II) ion forms if t	the reaction occurs.)_	
4) Zin	c chloride is synthesized from its eleme	nts	91 4		ž* –
5) Can	idle wax (C <sub>25</sub> H <sub>52</sub> ) is burned.				
6) Pho	sphoric acid and potassium hydroxide s	olutions are con	bined.		
7) Iron	(III) chloride is poured onto a piece of m	nagnesium		<del>.</del>	
8) Calciu	ım fluoride decomposes into its elemen	ts			
9) In a w	ell plate, magnesium chloride and alum	inum nitrate sol	utions are combined		-
10) Aqueo	us solutions of lead (II) nitrate and sodic	ım chloride are <sub>l</sub>	ooured together		

Chemical Reactions	Practice	Quiz	#2	Key
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	Per.	
Name_	 	_

Write the balanced equations for each of the following if a reaction occurs. If there is no reaction, write NR. Write the net ionic equation for all double displacement reactions that occur. Write the types of reaction in the space provided. Be sure to include states of matter.

1) Solutions of potassium carbonate and nitric acid (HNO₃) are combined. \_\_\_DD\_\_\_\_\_

 $K_2CO_{3(aq)} + 2 HNO_{3(aq)} \rightarrow 2 KNO_{3(aq)} + CO_{2(g)} + H_2O(I)$ 

(carbon dioxide and water from hydrogen carbonate)

2) Sodium metal is placed in water.\_\_SD\_\_\_\_\_

 $2Na(s) + 2H_2O(l) \rightarrow 2NaOH(aq) + H_{2(g)}$ 

3) Copper is placed in a solution of nickel (II) nitrate. (The copper (II) ion forms if the reaction occurs.)\_\_\_\_\_\_

NR

4) Zinc chloride is synthesized from its elements.\_\_\_\_synthesis\_\_\_\_\_

 $Zn(s) + Cl_{2(g)} \rightarrow ZnCl_{2}(s)$ 

5) Candle wax (C<sub>25</sub>H<sub>52</sub>) is burned. \_\_\_\_combustion\_\_\_\_

 $C_{25}H_{52(s)} + 38 O_{2(g)} \rightarrow 25 CO_{2(g)} + 26 H_2O(g)$ 

 $H^{+}(aq) + OH^{-}(aq) \rightarrow HOH(I)$ 

7) Iron (III) chloride is poured onto a piece of magnesium.\_\_\_\_SD\_\_\_\_\_

3 Mg(s) + 2 FeCl<sub>3(aq)</sub>  $\rightarrow$  2 Fe (s) + 3 MgCl<sub>2(aq)</sub>

8) Sodium fluoride decomposes into its elements.\_\_\_\_\_decomposition\_\_\_\_\_

 $2 \text{ NaF(s)} \rightarrow 2 \text{ Na(s)} + F_{2(g)}$ 

9) In a well plate, magnesium chloride and aluminum nitrate solutions are combined.\_\_\_\_DD\_\_\_\_

NR

10) Aqueous solutions of lead (II) nitrate and sodium chloride are poured together.\_\_\_\_DD\_\_\_\_

 $Pb(NO_3)_{2(aq)} + 2 NaCl(aq) \rightarrow 2 NaNO_{3(aq)} + PbCl_{2(s)}$ 

$$Pb^{2+}_{(aq)} + 2 Cl_{(aq)} \rightarrow PbCl_{2(s)}$$