

Chemical Reactions Practice Quiz #2

Name \_\_\_\_\_ Per. \_\_\_\_\_

Write the balanced equations for each of the following if a reaction occurs. If there is no reaction, write NR.  
Write the net ionic equation for all double displacement reactions that occur. Write the types of reaction in the space provided. Be sure to include states of matter.

- 1) Solutions of potassium carbonate and nitric acid ( $\text{HNO}_3$ ) are combined. \_\_\_\_\_
- 2) Sodium metal is placed in water. \_\_\_\_\_
- 3) Copper is placed in a solution of nickel (II) nitrate. (The copper (II) ion forms if the reaction occurs.) \_\_\_\_\_
- 4) Zinc chloride is synthesized from its elements. \_\_\_\_\_
- 5) Candle wax ( $\text{C}_{25}\text{H}_{52}$ ) is burned. \_\_\_\_\_
- 6) Phosphoric acid and potassium hydroxide solutions are combined. \_\_\_\_\_
- 7) Iron (III) chloride is poured onto a piece of magnesium. \_\_\_\_\_
- 8) Calcium fluoride decomposes into its elements. \_\_\_\_\_
- 9) In a well plate, magnesium chloride and aluminum nitrate solutions are combined. \_\_\_\_\_
- 10) Aqueous solutions of lead (II) nitrate and sodium chloride are poured together. \_\_\_\_\_

Write the balanced equations for each of the following if a reaction occurs. If there is no reaction, write NR.  
Write the net ionic equation for all double displacement reactions that occur. Write the types of reaction in the space provided. Be sure to include states of matter.

- 1) Solutions of potassium carbonate and nitric acid ( $\text{HNO}_3$ ) are combined. DD \_\_\_\_\_



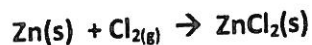
- 2) Sodium metal is placed in water. SD \_\_\_\_\_



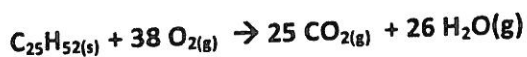
- 3) Copper is placed in a solution of nickel (II) nitrate. (The copper (II) ion forms if the reaction occurs.) \_\_\_\_\_

NR

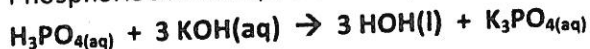
- 4) Zinc chloride is synthesized from its elements. synthesis \_\_\_\_\_



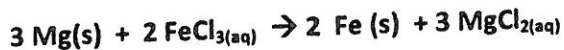
- 5) Candle wax ( $\text{C}_{25}\text{H}_{52}$ ) is burned. combustion \_\_\_\_\_



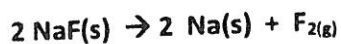
- 6) Phosphoric acid and potassium hydroxide solutions are combined. DD \_\_\_\_\_



- 7) Iron (III) chloride is poured onto a piece of magnesium. SD \_\_\_\_\_



- 8) Sodium fluoride decomposes into its elements. decomposition \_\_\_\_\_



- 9) In a well plate, magnesium chloride and aluminum nitrate solutions are combined. DD \_\_\_\_\_

NR

- 10) Aqueous solutions of lead (II) nitrate and sodium chloride are poured together. DD \_\_\_\_\_

